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An Arrhenius acid is any substance that produces hydrogen ions [or hydronium ions (H_3O^+) a hydrogen ion attached to a water molecule] when dissolved in water.

An Arrhenius base is any substance that produces hydroxide ions when dissolved in water. While the Arrhenius definitions of acids and bases is useful, it is limited.

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Acids and bases are important substances in many settings. In everyday life we encounter acids. such as vinegar and lemon juice and bases such as ammonia and lye in drain cleaners. Many. important processes in chemistry and biology involve reactions of acids with bases. Acid-base.

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In some textbooks, when explaining the Brønsted — Lowry definition, acids and bases are described as proton donors and proton acceptors. Based on your response to Exercise 2, explain why these are correct terms. ion desca\oed (1B 4. A definition of the prefix amphi is; "both or of both kinds."

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Mrs. Zuberbuehler - Mrs. Zuberbuehler
STOP Conjugate Acid $\text{CO}_2(\text{aq}) + \text{H}_2\text{O}(\text{aq})$ Model 3 — Conjugate Acid—Base
Pairs Base + 1-120 (l) HC031-(aq) 13. All
acid—base reactions have two conjugate
acid—base pairs. One conjugate acid—base
pair in the reaction in Model 3 is $\text{H}_2\text{O}^+/\text{H}_2\text{O}$.
List the other acid—base pair in the
reaction. 14.

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Phosphoric acid (H_3PO_4) is a triprotic
acid, but phosphorous acid (H_3PO_3) is a
diprotic acid, and hypophosphorous acid

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(H₃PO₂) is a monoprotic acid. a. Draw Lewis dot-diagrams for each of these three acids. All of the structures include a double bonded oxygen.

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